

Class 2a

More Equilibrium Concepts

A. Review

- Kinetic view of equilibrium: forward rate = backward rate

$$K_{eq} = \frac{k_1}{k_{-1}} = \left[\prod_i C_i^{\nu_i} \right]_{at\ eq}$$
- Thermodynamic view of equilibrium
 - $\Delta G^\circ_{rxn} = -RT \ln K_{eq}$
 - ΔG° and K_{eq} independent of pressure
 - $G = H - TS$; $dG = -SdT + VdP$
 - $K_{eq} = \prod_i a_i^{\nu_i}$
 - a_i = activity coefficient = $\frac{f_i}{f_i^\circ}$ [= P_i for ideal gases]
 - Table available for ΔG° ; $\Delta G_{react} = \sum_i \alpha_i (\Delta G^\circ)_i$
 - $K_{eq} = K_y K_{P_{tot}} = K_P = \frac{K_C}{K_T}$
 both K_y and $K_{P_{tot}}$ are dependent on total pressure if $\Delta n \neq 0$, but they compensate for each other so K_{eq} is independent of pressure!
 - Given K_{eq} and P_{tot} , we can find y_i 's @ equilibrium if we have y_{feed} (one reaction only)

B. More Concepts on equilibrium

- Hand-written notes...

C. Code Inputs

- Z_{1-7}^L low T range coefficients 300-1000 K
- Z_{1-7}^H high T range coefficients 1000-5000 K
- $$\frac{h^o}{RT} = \left(\sum_{i=1}^5 \frac{z_i T^{i-1}}{i} \right) + \frac{z_6}{T}$$
- $$\frac{S^o}{R} = z_1 \ln T + \left(\sum_{i=2}^5 \frac{z_i T^{i-1}}{i-1} \right) + z_7$$
- for ideal gas, $C_p - C_v = R$
- $dH = C_p dT$
(i.e., no pressure dependence for ideal gas)
- $dS = C_p d \ln T - R d \ln P$
(also, $dH = TdS + VdP$)
- $dU = C_v dT$

Approach

- Given expression for $H_i^\circ(T)$ and $S_i^\circ(T)$
 Get S_i $S_i - S_i^0 = \int_{P_0}^P -R d \ln P$
- Get S and H $S = \sum_i n_i S_i$
 $H = \sum_i n_i H_i$
- get $G = H - TS$
 minimize G

Sample Input Thermo Data

- (from NASA-Lewis)

Name	Source & Date	Elemental Composition	Phase	T_{low}	T_{high}	Card Number
H2	J 3/61H 200 000 000 0G	300.000 5000.000				1
	0.30989658E 01 0.51321595E-03 0.51397641E-07-0.34609047E-10 0.36695462E-14					2
	-0.87694041E 03-0.19569318E 01 0.30564452E 01 0.26833104E-02-0.58288567E-05					3
	0.55426526E-08-0.18209110E-11-0.98878140E 03-0.22961187E 01					4
Z%, low T range		Z%, high T range				